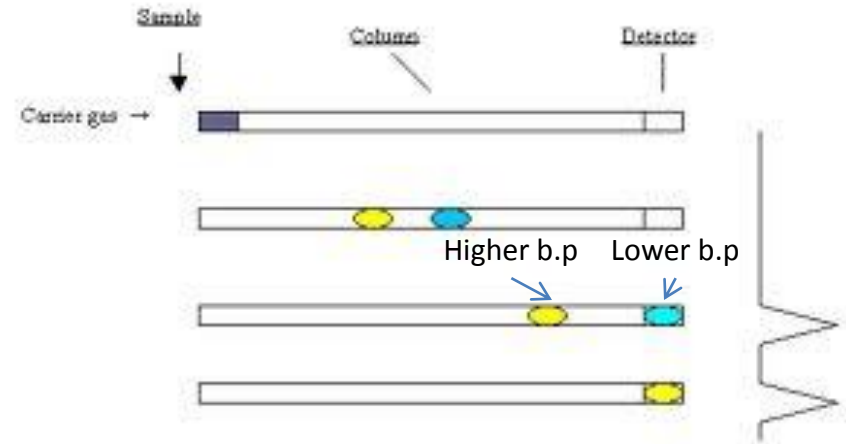
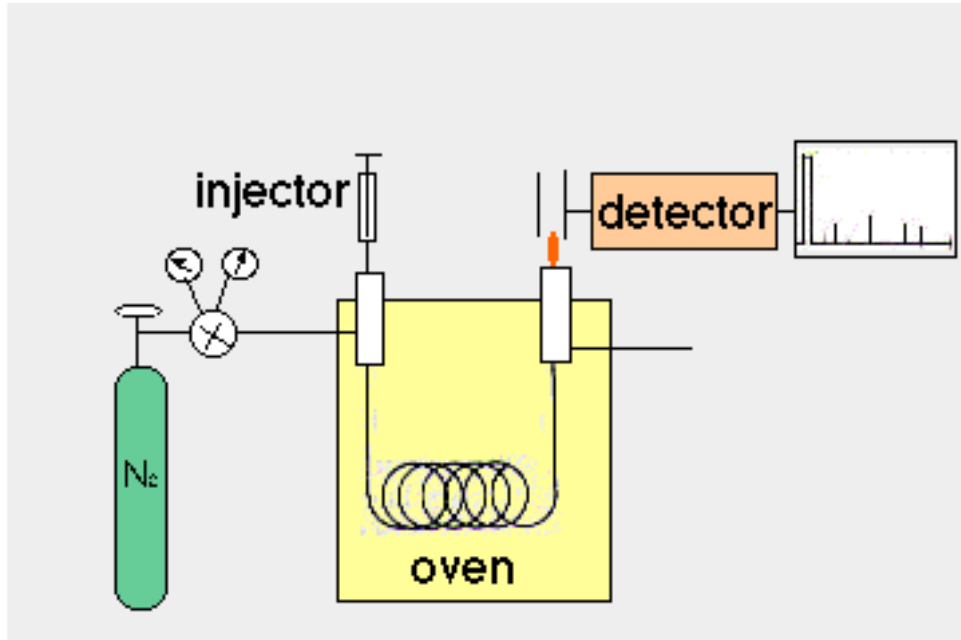
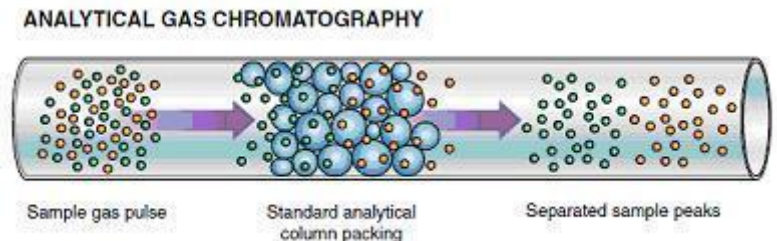


# Chapter 22 - Gas Chromatography

Column separation (gas-liquid, gas-solid) used for separating and analyzing compounds that can be vaporized without decomposition.



Separations based on differences in boiling points!



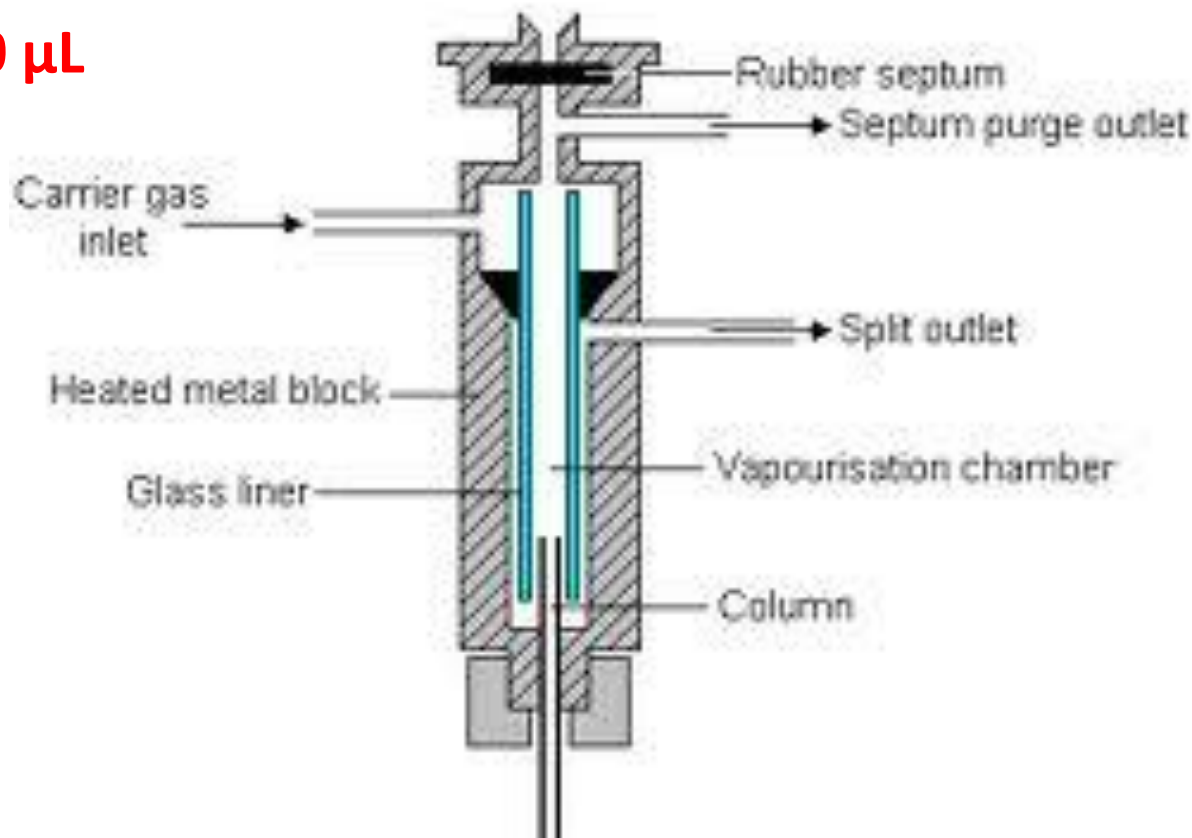
# Conditions

- Carrier gas (mobile phase) does NOTHING in GC but transport the compounds. Not involved in separation mechanism ( $H_2$  and He common).
- Injection volume (0.1 – 10  $\mu\text{L}$  generally). Temperature of injector is 50  $^\circ\text{C}$  greater than least volatile (highest boiling point compound). All compounds must be vaporized before transport onto column.
- Carrier gas is often dried by passage over molecular sieves as they strongly retain water. Activated by heating to 300  $^\circ\text{C}$  in vacuum.
- Gaseous mobile phase carries gaseous compounds (analytes) through a long column with a stationary phase.

# Injector

The split / splitless injector

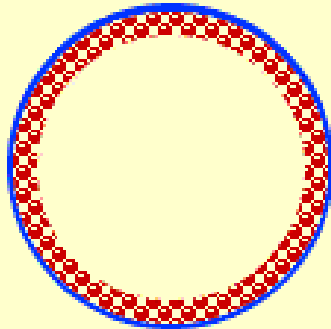
0.1 to 10  $\mu\text{L}$



Must vaporize all of the analytes in sample. Often the injector temperature is about 50 °C the least volatile analyte.

# Column Types

**Porous Layer Open  
Tubular Column.  
(PLOT) Column**



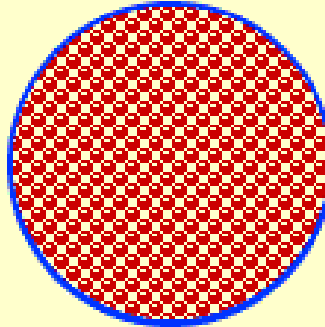
**Particle Layer  
Thickness**

**5-50  $\mu\text{m}$**

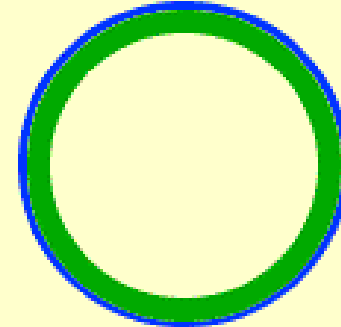
**Tube I.D.**

**320-530  $\mu\text{m}$**

**Packed Capillary  
Column  
(I.D.<1mm)**



**Wall Coated Open  
Tubular Column  
(WCOT) Column**



**Film  
Thickness**

**0.1-0.8  $\mu\text{m}$**

**Tube I.D.**

**100-530  $\mu\text{m}$**

$$N = L/H$$

*van Deemter equation  
for plate height:*

$$H \approx A + \frac{B}{u} + Cu$$

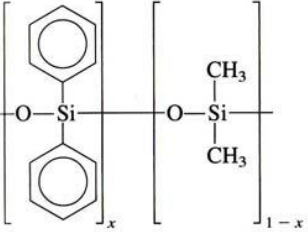
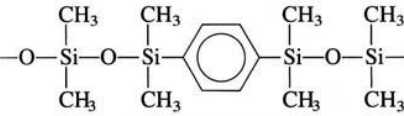
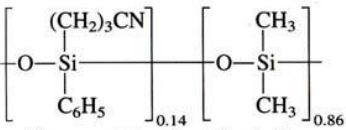
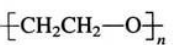
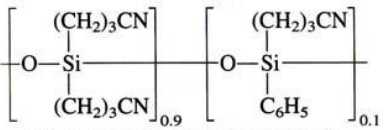
Multiple  
paths

Longitudinal  
diffusion

Equilibration  
time

# Stationary Phase Chemistry

**Table 22-1** Common stationary phases in capillary gas chromatography

Structure	Polarity	Temperature range
 <p>(Diphenyl)<sub>x</sub>(dimethyl)<sub>1-x</sub> polysiloxane</p>	<p><math>x = 0</math> Nonpolar  <math>x = 0.05</math> Nonpolar  <math>x = 0.35</math> Intermediate polarity  <math>x = 0.65</math> Intermediate polarity</p>	<p><math>-60^\circ</math> to <math>360^\circ\text{C}</math>  <math>-60^\circ</math> to <math>360^\circ\text{C}</math>  <math>0^\circ</math> to <math>300^\circ\text{C}</math>  <math>50^\circ</math> to <math>370^\circ\text{C}</math></p>
 <p>Arylene polysiloxane</p>	<p>Arylene stationary phases with low "bleed" (less thermal decomposition) at high temperature are available in compositions similar to other polysiloxanes in this table.</p>	
 <p>(Cyanopropylphenyl)<sub>0.14</sub>(dimethyl)<sub>0.86</sub> polysiloxane</p>	<p>Intermediate polarity</p>	<p><math>-20^\circ</math> to <math>280^\circ\text{C}</math></p>
 <p>Carbowax (polyethylene glycol)</p>	<p>Strongly polar</p>	<p><math>40^\circ</math> to <math>250^\circ\text{C}</math></p>
 <p>(Biscyanopropyl)<sub>0.9</sub>(cyanopropylphenyl)<sub>0.1</sub> polysiloxane</p>	<p>Strongly polar</p>	<p><math>0^\circ</math> to <math>275^\circ\text{C}</math></p>

# Separation Mechanisms

Different compounds have different *retention times*. For a particular compound, the retention time will vary depending on:

**The boiling point of the compound.** A compound which boils at a temperature higher than the column temperature is going to spend nearly all of its time condensed as a liquid at the beginning of the column. So high boiling point means a long retention time.

**The solubility in the liquid phase.** The more soluble a compound is in the liquid phase, the less time it will spend being carried along by the gas. High solubility in the liquid phase means a high retention time.

**The temperature of the column.** A higher temperature will tend to excite molecules into the gas phase - because they evaporate more readily.

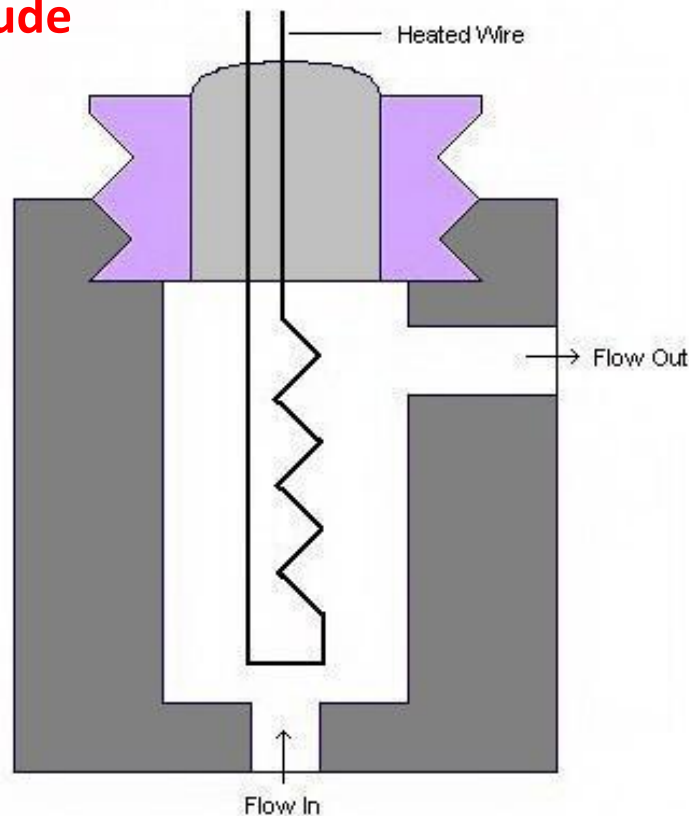
# Ideal Detector Characteristics

- Adequate sensitivity
- Good reproducibility and stability
- Dynamic range of 3-5 orders of magnitude
- Fast response time
- Stable to temperatures of 400 °C
- Nondestructive to the sample

# Thermal Conductivity Detector

LDR = 3-5 orders of magnitude

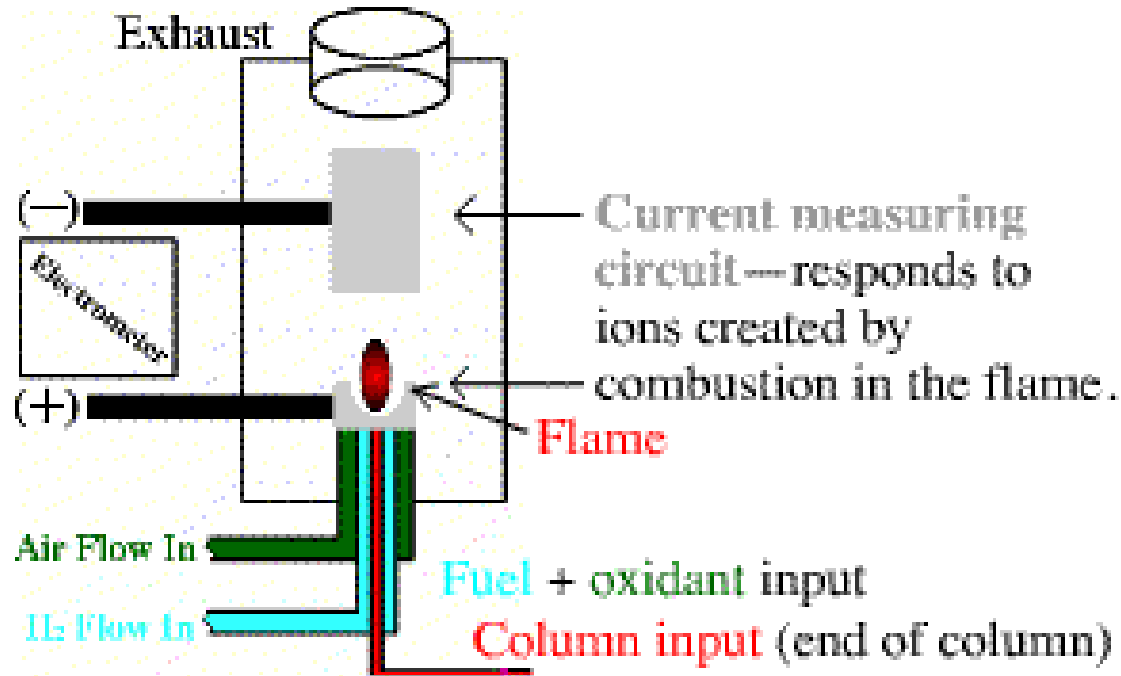
LOD = low ppm ( $S/N > 3$ )



**Universal detector! Thermal conductivity of gas decreases with analyte present.**



# Flame Ionization Detector (FID)



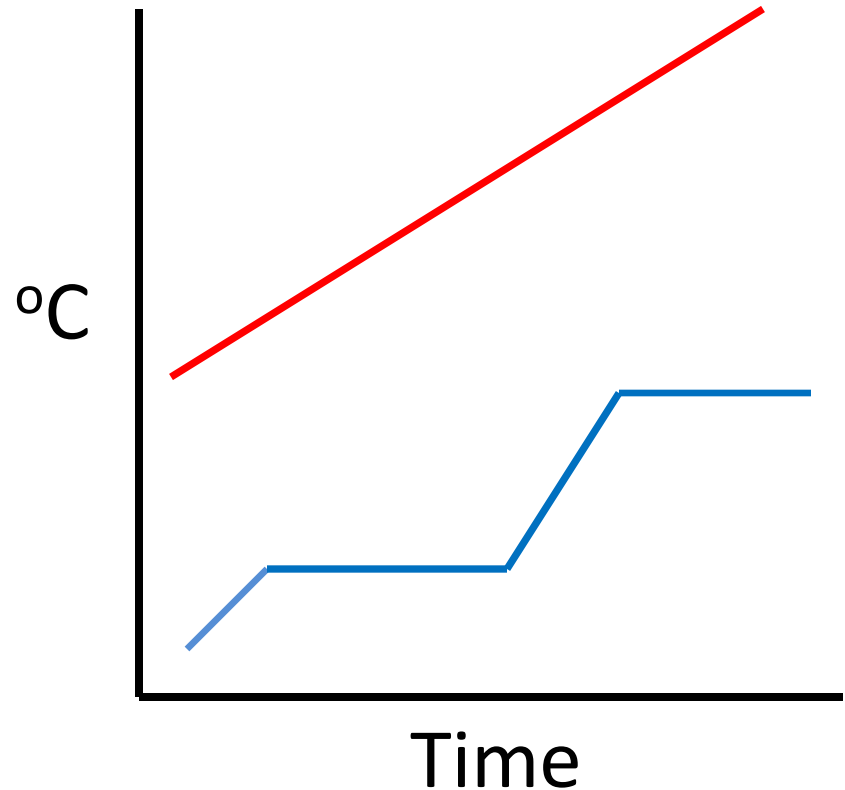
**Excellent for organic molecules! Not responsive to H<sub>2</sub>O, CO<sub>2</sub>, SO<sub>2</sub>, NO<sub>x</sub>.**

**LDR = 5-7 orders of magnitude**

**LOD = low ppb (S/N>3)**

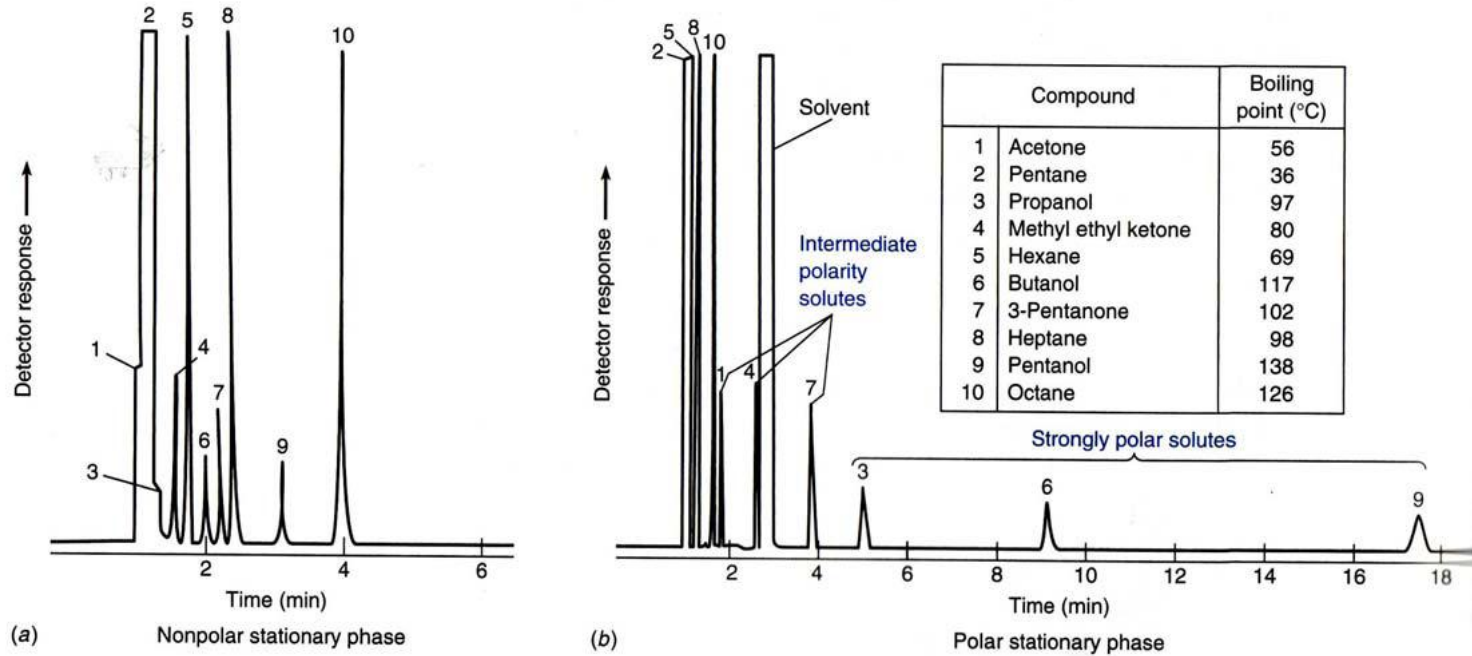
# Temperature Programming

- Raising column temp. decreases retention times.
- Sharpens peaks
- **Necessary** for separating compounds with a wide range of boiling points or polarities



**Fixed temperature separation = *isothermal* separation**

# Real Separation



**Figure 22-3** Separation of compounds on (a) nonpolar poly(dimethylsiloxane) and (b) strongly polar polyethylene glycol stationary phases (1  $\mu\text{m}$  thick) in open tubular columns (0.32 mm diameter  $\times$  30 m long) at 70°C. [Courtesy Restek Co., Bellefonte, PA.]

# Real Separation

