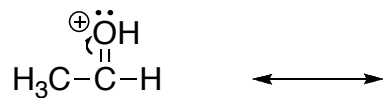
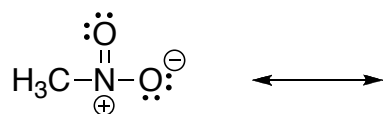
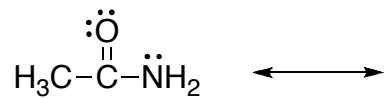
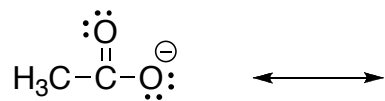


**CEM143, Problem Set #1**  
**Chapters 1-2**

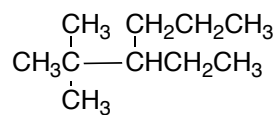
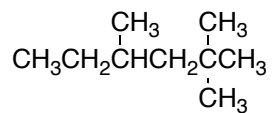
1. Draw Lewis dot structures and stick structures of the following compounds and calculate the formal charge for the indicated atoms:

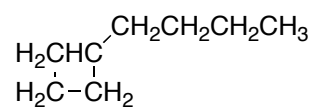
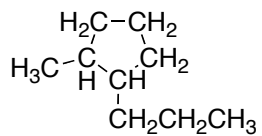
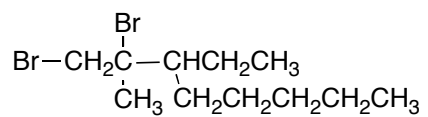
	Lewis Structures	Stick Structures	Formal Charges
CH <sub>3</sub> I			C =
NH <sub>3</sub>			N =
NH <sub>4</sub> <sup>⊕</sup>			N =
CH <sub>3</sub> NO <sub>2</sub>			N =
CH <sub>2</sub> N <sub>2</sub>			C =

2. Draw a resonance structure for the following compounds:



3. Convert the following structural (stick) formulas into line formulas and name the compounds:





4. Draw the structures that correspond to the following names:

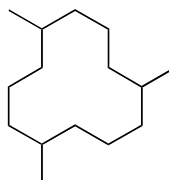
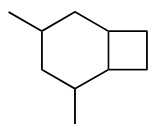
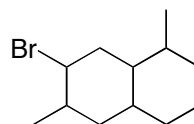
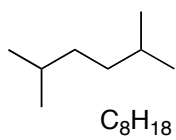
butyl bromide

t-butyl chloride

iso-propyl chloride

sec-butyl iodide

5. Provide the chemical formulas for:



6. Draw all possible Tribromopropanes: