A process for producing organic substituted aromatic or heteroaromatic compounds including biaryl and biheteroaryl compounds in a two-step reaction. In the first step, the aromatic or heteroaromatic compound is borylated in a reaction comprising a borane or diborane reagent (any boron reagent where the boron reagent contains a B—H, B—B or B—Si bond) and an iridium or rhodium catalytic complex. In the second step, a metal catalyst catalyzes the formation of the organic substituted aromatic or heteroaromatic compound from the borylated compound and an electrophile such as an aryl or organic halide, trflate (OSO₂CF₃), or nonaflate (OSO₂C₆F₁₃). The steps in the process can be performed in a single reaction vessel or in separate reaction vessels. The present invention also provides a process for synthesis of complex polyphenylenes starting from halogenated aromatic compounds.

24 Claims, 4 Drawing Sheets
FIGURE 2
FIGURE 3A

1. HBPin, 2 mol% 3, 2 mol% dppe, 100 °C, 16 h
2. 3-bromotoluene, K$_3$PO$_4$, 2 mol% Pd(PPh$_3$)$_4$, DME, 80 °C, 17 h

80% yield from 1,2-dichlorobenzene

FIGURE 3B

1. HBPin, 2 mol% 3, 2 mol% dppe, 100 °C, 16 h
2. K$_3$PO$_4$, 0.5 mol% Pd$_2$(dba)$_3$, 1.5 mol% Bu$_3$P, refluxing xylene, 15 h
\[ (PR_3)_n Ir^{III} (BPin)(E)_2 \]

\[ \text{C–H activating species} \]

\[ (PR_3)_n Ir^{II} (BPin)(E)_4 \]

\[ \text{H–BPin} + E \rightarrow E–E \]

\[ \text{E = H, BPin} \]

\[ n = 1, 2 \]

\[ \text{Fig. 4} \]
PROCESS FOR THE CATALYTIC SYNTHESIS OF BIARYLS AND POLYMERS FROM ARYL COMPOUNDS

CROSS-REFERENCE TO RELATED APPLICATION

This application claims priority to Provisional Application No. 60/305,107, filed Jul. 13, 2001, and to Provisional Application No. 60/332,092, filed Nov. 21, 2001.

STATEMENT REGARDING FEDERALLY SPONSORED RESEARCH OR DEVELOPMENT

This invention was supported in part by National Institutes of Health, National Institute of General Medical Sciences Grant No. R01 GM63188-01 and in part by National Science Foundation Grant No. CHE-9817290. The U.S. government has certain rights in this invention.

REFERENCE TO A “COMPUTER LISTING APPENDIX SUBMITTED ON A COMPACT DISC”

Not Applicable.

BACKGROUND OF THE INVENTION

(1) Field of the Invention

The present invention relates to a process for producing organic substituted aromatic or heteroaromatic compounds including biaryl and biheteroaryl compounds in a two-step reaction. In the first step, the aromatic or heteroaromatic compound is borylated in a reaction comprising a borane or diborane reagent (any boron reagent where the boron reagent contains a B-H, B-B or B-Si bond) and an iridium or rhodium catalytic complex. In the second step, a metal catalyst catalyzes the formation of the organic substituted aromatic or heteroaromatic compound from the borylated compound and an electrophile such as an organic or aroyl halide, triflate (OSO₂CF₃), or nonaflate (OSO₂CF₃). The steps in the process can be performed in a single reaction vessel or in separate reaction vessels. The present invention also provides a process for synthesis of complex polyphenylenes starting from halogenated aromatic compounds.

(2) Description of Related Art

Carbon-carbon bonds are the molecular “bricks and mortar” from which diverse architectures in living organisms and man-made materials are constructed. As the field of organic chemistry has evolved, numerous methods for carbon-carbon bond construction have been developed, ranging from classic examples, like the Diels-Alder reaction, to more recent metal-catalyzed processes such as olefin polymerizations and metatheses.


Arylboration reagents are typically synthesized in a multi-step process such as that shown below.

![Arylboration Reaction](attachment://arylboration.png)

Shorter routes that avoid undesirable halogenated aromatic intermediates would be attractive. Towards this end, theoretical estimates of B—H and B—C bond enthalpies gave credence to organoborate synthesis via the thermal dehydrogenative coupling of B—H and C—H bonds as shown below (Raben et al., J. Am. Chem. Soc. 116: 4121–4122 (1994)).

![Dehydrogenative Coupling](attachment://dehydrogenative.png)

Some key steps in putative catalytic cycles for this process had been established with Hartwig’s (Waltz et al., J. Am. Chem. Soc. 117: 11357–11358 (1995); Waltz and Hartwig, Science 277: 211–213 (1997)) and Marder’s reports of stoichiometric borylations of arenes, alkenes, and alkynes by metal boryl complexes (M-BR₂). Although arene activation products were not mentioned, small peaks in the GC-MS trace with masses consistent with toluene borylation products were assigned in the Supplementary Material to Nguyen et al., Am. Chem. Soc. 115, 9329–9330 (1993).

While Hartwig has developed elegant photochemical methods for hydrocarbon borylation using catalytic amounts of metal complexes (Chen and Hartwig, Angew. Chem. Int. Ed. 38: 3391–3393 (1999)), thermal, catalytic borylations of unactivated hydrocarbons had not been documented prior to our report in 1999 (Iverson and Smith, III, J. Am. Chem. Soc. 121: 7696–7697 (1999)). Since then, borylation of aliphatic and alkyl branched aliphobic hydrocarbons at a primary C—H hydrocarbon bond under thermal conditions using a rhodium catalytic complex which includes an electron donor ligand was disclosed in WO 01/64889 A1 and
of a metal catalyst wherein the aromatic group of the organic compound is coupled to the aromatic group of the aromatic boron compound to produce the organic substituted aryl or heteroaryl compound.

The present invention further provides a process for producing a polyphenylene, which comprises (a) reacting a mixture of aromatic compounds with one to five halogen groups and a borane in the presence of a catalytically effective amount of an iridium or rhodium complex with three or more substituents, and a phosphorus, carbon, nitrogen, oxygen, or sulfur organic ligand to form a mixture of borylated aromatic compounds; and (b) reacting the mixture of borylated aromatic compounds in the presence of a catalytically effective amount of a metal catalyst wherein the borylated aromatic compounds in the mixture are cross-coupled to produce the polyphenylene.

The present invention further provides process for producing a polyphenylene, which comprises (a) reacting a mixture of aromatic compounds and a borane in the presence of a catalytically effective amount of an iridium or rhodium complex with three or more substituents, and a phosphorus, carbon, nitrogen, oxygen, or sulfur organic ligand to form a mixture of borylated aromatic compounds; and (b) reacting the mixture of borylated aromatic compounds with a mixture of halogenated aromatic compounds with at least two halogen groups in the presence of a catalytically effective amount of a metal catalyst wherein the borylated aromatic compounds in the mixture are cross-coupled to the halogenated aromatic compounds to produce the polyphenylene.

In a further embodiment of the above processes, the three or more substituents excludes hydrogen.

In a further embodiment of the above processes, the iridium complex is selected from the group consisting of (Cp*)IrH₂(Me₃P), (Cp*)IrH(BPin)(Me₃P), (Cp*)IrH₂(C₂H₅)(Me₃P), (Ind)Ir(COD), (Ind)Ir(dppe), (MesH)Ir(BPin)(B(OR)₂)₂, ((R₂₃P₃)Ir(B(OR)₃)₂, (R₂₃P₃)Ir(BPin)₃, (R₂₃P₃)Ir[(R(O)₂)B(OR)₃]₂, (R₂₃P₃)Ir[(BPin)₃], (MesH)Ir(BPin)₃, and (IrCl(COD))₃, (Me₃P)IrH₂, (R₂₃P₃)IrH₂, and (R₂₃P₃)Ir[(B(OR)₃)₂]ₓ, where x is 0-4, wherein Cp* is 1,2,3,4,5-methylcyclopentadienyl, BPin is pinacolborane, Me is methyl, H is hydrogen, P is phosphorus, Ind is indenyl, COD is 1,5-cyclooctadiene, MesH is mesitylene, and wherein R, R₁, and R₂ are hydrogen, linear or branched alkyl containing 1-8 carbons, aryl, or a carbon in a cyclic structure.

In a further embodiment of the above processes, the iridium complex is (Ind)Ir(COD) wherein Ind is indenyl and COD is 1,5-cyclooctadiene, (MesH)Ir(BPin)₃, wherein MesH is mesitylene and BPin is pinacolborane, or (IrCl(COD))₃, wherein COD is 1,5-cyclooctadiene.

In a further embodiment of the above processes, the rhodium complex is selected from the group consisting of (Cp*)RhH₂(Me₃P), (Cp*)RhH(BPin)(Me₃P), (Cp*)RhH(C₂H₅)(Me₃P), and (Cp*)Rh(hexamethylbenzene), wherein Cp* is 1,2,3,4,5-methylcyclopentadienyl, BPin is pinacolborane, Me is methyl, H is hydrogen, and P is phosphorus.

In a further embodiment of the above processes, the phosphorus organic ligand is selected from the group consisting of trimethyl phosphine (PM₃), 1,2-bis(dimethylphosphino)ethane (dmpe), and 1,2-bis(diphenylphosphino)ethane (dppe).

In a further embodiment of the above processes, the borane is a borane ester.

In a further embodiment of the above processes, the borane is pinacolborane. In a further embodiment of the above processes, the metal is palladium.
In a further embodiment of the above processes, the metal catalyst complex is selected from Pd(PPh₃)₄, Pd₂(dbq)₃/P (Bu₄N)₂PdCl₂(dppf), Pd(OC₂H₅)₂(PC₅), wherein P is phosphorus and Ph is phenyl, dbq is dibenzylideneacetonate, Bu is tert-butyl, dppf is diphenylphosphinoferrocene.

OBJECTS

It is an object of the present invention to provide a process for producing organic substituted aryl compounds and polyphenylenes.

It is a further object of the present invention to provide a two-step process for producing organic substituted aryl compounds and polyphenylenes.

It is a further object of the present invention to provide a two-step process that can be carried out in a single reaction vessel.

These and other objects of the present invention will become increasingly apparent with reference to the following drawings and preferred embodiments.

DETAILED DESCRIPTION OF THE DRAWINGS

FIG. 1 shows the formulas for precatalysts 1 to 15. CP is 1,2,3,4,5-methylcyclopentadienyl, BP is pinacoloborane, Me is methyl, H is hydrogen, P is phosphorus, Ind is indenyl, COD is 1,5-cyclooctadiene, MesH is mesitylene, and wherein R₁, R₂, and R₃ are each selected from the group consisting of hydrogen, linear or branched alky containing 1 to 8 carbons, aryl, and a carbon in a cyclic structure.

FIG. 2 shows the formulas for precatalysts 16 to 27. Y₁, and Y₆ are each selected from the group consisting of hydrogen, halide, alkyl, aryl, alkoxide (—O(R₃)₂), and amide (—N(R₃)₂(R₄)) wherein R₁, R₁₂, and R₁₃ are each selected from the group consisting of hydrogen, linear alky containing 1 to 8 carbon atoms, branched alky containing 1 to 8 carbons, and a carbon in a cyclic structure; R₁₄, R₁₅, and R₁₆ are each selected from the group consisting of hydrogen, linear alky, branched alky, and a carbon in a cyclic structure; (P₇)P is R₁₇R₁₈, and R₁₉, and Y₇ is a chain containing 1 to 12 carbons; (P₇)P is of the formula

wherein R₂₀, R₂₁, R₂₂, R₂₃, R₂₄, R₂₅, R₂₆, and R₂₇ are each selected from the group consisting of alkyl chains, carbocyclic rings, and aryl groups; and BY is a boron moiety.

FIG. 3A shows an example of the one vessel C—H activation/cross coupling reactions as applied to biaryl synthesis.

FIG. 3B shows an example of the one vessel C—H activation/cross coupling reactions as applied to polyphenylene synthesis.

FIG. 4 shows a mechanism for aromatic borylations catalyzed by Ir boryl complexes.

DETAILED DESCRIPTION OF THE INVENTION

All patents, patent applications, provisional patent applications, government publications, government regulations, and literature references cited in this specification are hereby incorporated herein by reference in their entirety. In case of conflict, the present description, including definitions, will control.

For convenience, wherever the term “aromatic” is used alone, it is to be construed to include both aromatic and heteroaromatic compounds selected from the group consisting of aryl, six membered heteroaromatic compounds, and five membered heteroaromatic compounds and the wherever the term “organic substituted aromatic compounds” is used alone, it is to be construed to include both organic substituted aromatic and heteroaromatic compounds and biaryl compounds.

The present invention provides a process for producing organic substituted aromatic and heteroaromatic compounds including biaryl and heteroaryl compounds, which comprises in a first step, reacting an aromatic or heteroaromatic compound with a borane or diborane containing a B—H, B—Br or B—Si bond, but preferably a borane ester, in the presence of a catalytically effective amount of an iridium or rhodium complex with three or more substituents to produce a borylated aromatic or heteroaromatic intermediate; and in a second step, reacting the borylated aromatic or heteroaromatic intermediate with an organic electrophile such as an aryl halide, trflate (OSO₂CF₃), or nonflate (OSO₂CF₃)₂ or the like, in the presence of a catalytically effective amount of a metal catalyst wherein the aryl group of the aryl or organic halide, trflate, or nonflate is coupled to the aromatic or heteroaromatic group of the borylated compound to produce the organic substituted aromatic or heteroaromatic compound, including biaryl.

The borylation of an aromatic or heteroaromatic substrate and the subsequent coupling of the borylated aromatic or heteroaromatic intermediate with a second sp³-hybridized halocarbon can be performed in a single reaction vessel as shown in Scheme 1.

\[ \begin{align*}
\text{Ar} & \rightleftharpoons \text{H} + \text{H}(\text{OR})₂ \\
\text{Ar} &= \text{H} + \text{H}(\text{OR})₂ \\
\text{Ar} &= \text{H} + \text{H}(\text{OR})₂
\end{align*} \]

The present invention further provides a process for producing a polyphenylene, which in a first step comprises reacting a mixture of aromatic compounds with one to five halogen groups and a borane or diborane containing a B—H, B—Br or B—Si bond, but preferably a borane ester, in the presence of a catalytically effective amount of an iridium or rhodium complex with three or more substituents, to form a mixture of borylated aromatic intermediates; and in a second step, reacting the mixture of borylated aromatic intermediates in the presence of a catalytically effective amount of a metal catalyst wherein the borylated aromatic intermediates in the mixture are cross-coupled to produce the polyphenylene. See FIG. 3B for an example of the process. The mixture of aromatic compounds can contain all of one type of aromatic compound or a mixture of different types of aromatic compounds.
The present invention further provides a process for producing a polyphenylene, which in a first step comprises reacting a mixture of aromatic compounds and a borane or diborane containing a B–H, B–B, or B–Si bond, but preferably a borane ester, in the presence of a catalytically effective amount of an iridium or rhodium complex with three or more substituents, to form a mixture of borylated aromatic intermediates; and in a second step, reacting the mixture of borylated aromatic intermediates with a mixture of halogenated aromatics with at least two halogen groups in the presence of a catalytically effective amount of a metal catalyst wherein the borylated aromatic intermediates in the mixture are cross-coupled to the halogenated aromatic compounds to produce the polyphenylene. The mixture of aromatic compounds can contain all of one type of aromatic compound or a mixture of different types of aromatic compounds. The process is shown in Scheme 2 below.

![Scheme 2](image)

In Scheme 2, HB(OR)₂ is preferably BPin.

The preferred catalysts for producing the borylated intermediates in the first step of the process comprise iridium (Ir) or rhodium (Rh) in a complex with three or more substituents, preferably excluding hydrogen, bonded to the Ir or Rh and preferably, further including a phosphorus organic ligand, which is at least in part bonded to the Ir or Rh. The process for forming B–C bonds between boranes and sp²-hybridized C–H bonds to produce organoboron intermediates such as ring-substituted aranes (or aryl boron esters and acids) according is shown in Scheme 3.

![Scheme 3](image)

The direct route to aromatic or heteroaromatic boronate esters and acids in the first step produces intermediates which are versatile transfer reagents in the second step of the process of the present invention. The boron in these transfer reagents serves as a mask for a broad range of heteroatoms and functional groups during the catalytic cross-coupling reactions of C–B and C–X (X is Cl, Br, I, triflate, nonflate) groups in the second step to yield new C–C bonds as shown in scheme 4.

![Scheme 4](image)

The above two-step process is particularly useful in the pharmaceutical industry for drug manufacturing and for synthesis of compounds in drug discovery. Thus, the present invention further provides catalyst kits that can be used for general couplings in drug discovery applications.

Preferably, the B–C bond-forming reaction between a borane and an sp²-hybridized C–H bond to produce a ring substituted arene in the first step is catalyzed by a catalyst comprising Ir and Rh in a complex with three or more substituents, preferably excluding hydrogen as a substituent, bonded to the Ir or Rh and further preferably, an organic ligand selected from the group consisting of phosphorus, carbon, nitrogen, oxygen, and sulfur organic ligands. For example, phosphorus organic ligands, organic amines, imines, nitrogen heterocycles, ethers, and the like. Preferably, the ligand is in a molar ratio about 1 to 1 to 1, wherein the organic ligand is at least in part bonded to the iridium or rhodium.

Effective precatalysts for forming the B–C bonds can be grouped into two families: those that contain cyclopentadienyl (Cp⁺, Cp₂R₂ wherein R is CH₃) or indenyl (Ind, Cp₃R₇ wherein R is Ind ligands and those that contain phosphine ligands. Included are compounds that contain both the Cp⁺ and the Ind ligands and the phosphine ligands.

Preferably, the Ir catalytic composition for the first step of the process comprises one of the following: (ArH)Ir(BY)₃ wherein ArH is selected from the group consisting of aromatic, heteroaromatic, polyaromatic, and heteropolycyclic hydrocarbon and wherein BY is a boron moiety; (MesH)Ir(BY)₃ wherein MesH is mesitylene and wherein BY is a boron moiety; (P(Y₁,Y₂,Y₃))Ir(H₂)(BY)₃₋ₙ wherein Y₁, Y₂, and Y₃ are each selected from the group consisting of hydrogen, halide, alkyl, aryl, alkoxide (—O(R₁)), and amide (—N(R₂,R₃)) wherein R₁, R₂, and R₃ are each selected from the group consisting of hydrogen, linear alkyl containing 1 to 8 carbon atoms, branched alkyl containing 1 to 8 carbons, and a carbon in a cyclic structure, wherein n is 0, 1, or 2, and wherein BY is a boron moiety; (P(R₁,R₂)(R₃,R₄))Ir(H₂)(BY)₃₋ₙ wherein R₁,R₂, and R₃,R₄ are each selected from the group consisting of hydrogen, linear alkyl, branched alkyl, and a carbon in a cyclic structure, wherein n is 0, 1, or 2, and wherein BY is a boron moiety; (P(Y₁,Y₂,Y₃))Ir(H₂)(BY)₃₋ₙ wherein Y₁, Y₂, and Y₃ are as above, wherein R₁,R₂, and R₃ are selected from the group consisting of a linear alkyl containing 1 to 8 carbon atoms, branched alkyl containing 1 to 8 carbons, aryl, and a carbon in a cyclic structure, and wherein BY is a boron moiety; (P(R₁,R₂)(R₃,R₄))Ir(H₂)(BY)₃₋ₙ wherein R₁,R₂, R₃,R₄, and R₅ are as above; R₆ is as above, and wherein BY is a boron moiety; (P(Y₁,Y₂,Y₃))Ir(BY)ₙ₋ₖ wherein BY is a boron moiety, wherein (P,Y) is R₁,R₂,P—Y₁,Y₂—P—R₅,R₆ wherein R₁,R₂, R₅, and R₆ are each selected from the
group consisting of hydrogen, linear alkyl containing 1 to 8 carbon atoms, branched alkyl containing 1 to 8 carbons, and a carbon in a cyclic structure, and wherein Y₅ is a chain containing 1 to 12 carbons; (PY₅)²Ir(Y₅)(Y₅)Ir(BPin) (17) wherein BY is a boron moiety, wherein Y₅, Y₅, and Y₅ are as above, and wherein (PY₅)² is as above; (PY₅)²Ir(R₃)(R₃)(BY) (18) wherein BY is a boron moiety, wherein R₃, R₃, and R₃ are as above, wherein (PY₅)² is as above; {[(P⁻)²Ir(BY)]₂, (μ₂-(P⁻)²) (19) wherein BY is a boron moiety and wherein (P⁻)² is of the formula

wherein R₉, R₉, and R₉ are each selected from the group consisting of alkyl chains, carboxylic rings, and aryl groups; (P⁻)²Ir(R₉)(R₉)(BY) (20) wherein BY is a boron moiety, wherein Y₅, Y₅, and Y₅ are as above, and wherein (P⁻)² is as above; (P⁻)²Ir(R₉)(R₉)(BY) (21) wherein BY is a boron moiety, wherein R₉, R₉, and R₉ are as above, wherein (P⁻)² is as above; (PY₅)²Ir(BY) (22) wherein BY is a boron moiety, and wherein (PY₅)² is as above; (P⁻)²Ir(BY) (23) wherein BY is a boron moiety, and wherein (P⁻)² is as above; (PY₅)²Ir(BY) wherein Y₅, Y₅, and Y₅ are as above and BY is a boron moiety; (R₉)(R₉)(R₉)Ir(BY) (24) wherein BY is a boron moiety, wherein Y₅, Y₅, and Y₅ are as above, and wherein (PY₅)² is as above; (P⁻)²Ir(R₉)(R₉)(BY) (25) wherein BY is a boron moiety, wherein R₉, R₉, and R₉ are as above, and wherein (P⁻)² is as above; (PY₅)²Ir(BY) (26) wherein BY is a boron moiety, R₉, R₉, and R₉ are as above, and wherein (PY₅)² is as above; (P⁻)²Ir(R₉)(R₉)(BY) (27) wherein BY is a boron moiety, wherein R₉, R₉, and R₉ are as above, and wherein (P⁻)² is as above.

Examples of catalytic compositions comprising iridium include those selected from the group consisting of (Cp)²Ir(R₉)(R₉)(BY) (1), (Cp)²Ir(BPin)(C₆H₅) (2), (Cp)²Ir(BPin)(C₆H₅)(C₆H₅) (3), (Ind)Ir(COD) (8), (Mes)Ir(BPin)(BOR) (9), (R₉)²Ir(BPin)(BOR) (10), (R₉)²Ir(BPin) (11, (R₉)²Ir(BPin) (12), (R₉)²Ir(BPin) (13), (R₉)²Ir(BPin) (14), (Mes)Ir(BPin) (9 wherein BOR is BPin), IrCl(COD) (15) and IrCl(COD) (16), wherein CP is 1,2,3,4,5-methylcyclopentadienyl, BPin is pinacoloborane, Me is methyl, H is hydrogen, P is phosphorus, Ind is indenyl, COD is 1,5-cyclooctadiene, MesH is mesitylene, and wherein R, R, and R are each selected from the group consisting of hydrogen, linear or branched alkyl containing 1 to 8 carbons, aryl, and a carbon in a cyclic structure.

Preferably, the Rh catalytic composition for the first step comprises one of the following: (Cp)²Ir(R₉)(R₉)(BY) (28) wherein Y₅, Y₅, and Y₅ are as above, wherein n is 0 or 1, wherein BY is a boron moiety, and wherein Cp is of the formula

wherein R₉, R₉, and R₉ are each selected from the group consisting of hydrogen, alkyl chains, carboxylic rings, and aryl groups; and (Cp)²Ir(R₉)(R₉)(BY) (29) wherein R₉, R₉, and R₉ are as above; n is 0 or 1, wherein BY is a boron moiety; and wherein Cp is as above.

Examples of catalytic compositions comprising rhodium include those selected from the group consisting of (Cp)²Rh(H)(Me)(P) (4), (Cp)²Rh(BPin)(P) (5), (Cp)²Rh(H)(C₆H₅)(Me)(P) (6), and (Cp)²Rh(hexamethylbenzene) (7), wherein CP is 1,2,3,4,5-methylcyclopentadienyl, BPin is pinacoloborane, Me is methyl, H is hydrogen, and P is phosphorus.

In the above catalytic compositions, preferably the BY boron moiety selected from the group consisting of

wherein R₉, R₉, and R₉ are each selected from the group consisting of hydrogen, linear alkyl containing 1 to 8 carbon atoms, branched alkyl containing 1 to 8 carbons, and a carbon in a cyclic structure. Figs. 1 and 2 show the structures of precatalysts 1 to 15 and 16 to 27, respectively.

While the precatalysts can under particular reaction conditions catalyze the borylation of particular ring-substituted arenes, the reactions proceed more efficiently when an organic ligand such as phosphine ligands (phosphorus containing ligands) are included in the reaction mixture. The addition of phosphine ligands to the reaction generates active catalysts which can produce ring-substituted arene boranes (aryl boronate esters and acids) with low catalyst loading. The fact that phosphine-containing species can catalyze borylation is important because numerous phosphines are commercially available. Furthermore, the selectivities of the borylation can be altered as a function of the phosphine ligand that is used. Examples of phosphine ligands include, but are not limited to, trimethyl phosphine (PM₅), 1,2-bis(dimethylphosphino)ethane (dme), 1,2-bis(diphenylphosphino)ethane (dppe), Cy₂P, and Ph₂P.

For example, precatalyst 8 can be obtained in two high-yielding steps from the common iridium starting material, IrCl₃(H₂O). Precatalyst 9 can be prepared by reacting 8 with approximately 5 equivalents of pinacoloborane (HPin) in mesitylene solvent. It was discovered that commercially available precatalyst 15 will also catalyze borylations. While all of the precatalysts have similar activities for many substrates, borylations of particular arenes exhibit a remarkable precatalyst dependence.
In the absence of phosphine ligands, compound 8 catalyzes the borylation of benzene by HBPin, but relatively high catalyst loading and long reaction times are required to prepare PhBP in reasonable yields. At temperatures above 80°C, decomposition to Ir metal occurs, which halts catalysis. Compound 9 is not effective in catalysis without the addition of phosphine.

Addition of phosphine ligands to solutions of compound 8 and 9 generates active catalysts for the production of aryl boronic esters with low catalyst loading as illustrated for the examples in FIG. 1. The fact that phosphine-containing species can catalyze borylation is important because numerous phosphines are commercially available. Consequently, the selectivities can be altered as a function of the phosphine that is added.

Another virtue of the present invention is that a broad range of heteroatoms and functional groups are inert under borylation conditions as shown in Scheme 5.

Given that Grignard reagents react with several of these groups and Pd catalyzes the formation of ArBPin from ArBr and HBPin, the functional group tolerance for the Ir-catalyzed chemistry is remarkable. Under appropriate conditions, even iodobenzene can be borylated without iodide reduction. In this instance, no conversion was observed when using precatalyst 8, whereas precatalyst 9 gives the borylated products in 95% yield as shown in Scheme 6.

Therefore, the present substrate compatibility, which is already remarkably broad, is expected to expand with further improvements to the present invention.

For monosubstituted arenes, mixtures of meta and para borylated products are obtained. In contrast to the known Rh complexes that catalyze aromatic borylation, the meta:para ratio deviates significantly from 2:1. For most substrates, this ratio exceeds 3:1 and data for anisole are shown in Table 1.

### Table 1

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<th>Catalyst</th>
<th>Temp (°C)</th>
<th>Time (h)</th>
<th>o.m.p</th>
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<td>2 mol % 8/2 PMe₂</td>
<td>150</td>
<td>29</td>
<td>9:74:17</td>
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<tr>
<td>2</td>
<td>2 mol % 9/2 PMe₂</td>
<td>150</td>
<td>41</td>
<td>8:75:17</td>
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<td>22</td>
<td>3:76:21</td>
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<tr>
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<td>51</td>
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<td>100</td>
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<td>150</td>
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<td>12:53:36*</td>
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<tr>
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<td>150</td>
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<td>2:57:40</td>
</tr>
<tr>
<td>9</td>
<td>2 mol % 9/PMe₂</td>
<td>150</td>
<td>40</td>
<td>3:67:30</td>
</tr>
</tbody>
</table>

*Low conversion.

It is noteworthy that the para isomer is more favored for entries 7 and 8, where the meta:para ratio is significantly less than 2:1. These data show that while there is a steric bias against ortho borylation, the meta:para ratio is sensitive to the type and amount of phosphine ligands that are added. For dppe, the activity at 100°C is relatively high, and the reaction is complete in less time than at 150°C.

With the exception of F and amide functional groups in some Rh catalyzed reactions, borylation at positions that are ortho to functional groups are avoided. Thus, 1,3-substituted aromatics can be selectively borylated at the 5' position. This is the hardest position to selectively activate by traditional aromatic substitution chemistry and for electron rich arenes, there are no general methods for preparing derivatives from the 1,3-substituted arenes.

Furthermore, multiple borylation of 1,3-substituted arenes does not occur to a significant extent, which means that equimolar quantities of borane and arene give aromatic boronic esters in high yield in the absence of solvent. Substrates that have been successfully converted to boronate esters under these conditions are shown in scheme 7.

For fluorinated benzenes, the borylation at ortho positions occurs readily. Hence, C₆H₅F₃ and 1,3,5-trifluorobenzene give mono and triborylated products, respectively, as shown in Scheme 8.
It is noteworthy that present Rh catalysts are not compatible with halide functionalities and substantial quantities of dehalogenated and diborylated products are observed. We extended the chemistry to five-membered rings and heterocycles as shown by the borylation of a protected pyrrole and 2,6-lutidine in Scheme 9.

The borylation of aromatic and heteroaromatic compounds and the catalysts suitable for the borylation are disclosed in the commonly owned U.S. Application, which was filed Jul. 13, 2002, and which claims priority to U.S. Provisional 60/305,107 filed Jul. 13, 2001.

In the second step of the process of the present invention, the borylated intermediate is reacted with a halogenated aryl or organic, triflate, or nonaflate compound in the presence of a metal catalyst which cross-couples the C at the C—B bond and the C at the C-halogen bond to produce the substituted aromatic, heteroaromatic, biaryl, or bibeteroaryl compound. In a preferred embodiment, the metal comprising the metal catalyst is palladium. Examples of Pd catalysts which are suitable for the cross-coupling include, but are not limited to, Pd(PPh3)4/P(Bu)3, PdCl2(dppf) Pd(OAc)2/PCy3, wherein P is phosphorus and Ph is phenyl, dba is dibenzylacetylene, ‘Bu is tert-butyl, and dppf is diphenylphosphinophenylacetylene.

The process for making organic substituted aromatic and heteroaromatic compounds, including biaryls and bibeteroaryls, avoids many of the limitations of the prior art because (1) the borylation reactions can be carried out in neat substrates, thereby avoiding ethereal solvents, (2) since the C—H bonds are selectively activated, halogenation of arenes and conversions to Grignard or organolithium reagents are eliminated, (3) the only byproducts of the borylation reaction are hydrogen, which is easily removed, and the catalyst, which is present in low concentrations, can be recovered, (4) the process of the present invention tolerates a broad range of functional groups, (5) active catalysts are generated from common precursors and selectivities can be altered by adding commercially available ligands such as alkyl phosphines, (6) particular substitution patterns which are notoriously difficult to achieve using prior art aromatic substitution chemistry can be obtained in one step starting from inexpensive starting materials, and (8) Ir metal is relatively inert, Ir complexes generally have low toxicity, and Ir metal recovered from the reactions can be recovered from the reaction waste and recycled. Furthermore, the process can be used to make chiral organic substituted aromatic compounds.

The following examples are intended to promote a further understanding of the present invention.

**EXAMPLE 1**

The process of the present invention was inspired by Bergman’s (Science 223: 902–908 (1984)) and Jones and Fehrer’s (Acc. Chem. Res. 22: 91–100 (1989)) fundamental studies of hydrocarbon (R—H, R—silyl or aryl) activation by Cp*Me)(M’)(M=Ir, Rh; Cp* η-C5Me5), which produce Cp*(PMc)(M’)(H)(R) where M—H and M—R bonds result from R—H scission.

While investigating stoichiometric B—C bond formation reactions between Cp*(PMc)(Ir)(H)(Ph) and pinacolborane (HBPin), we found that substantial quantities of aryloborin products were produced from catalytic solvent activation. The major metal-containing product in this reaction, Cp*(PMc)Ir(H)(BPin)(2), was a precatalyst for benzene borylation with an effective turnover number (TON) corresponding to the formation of three molecules of PhBPin per molecule of 2 (Iverson and Smith, III, J. Am. Chem. Soc. 121: 7696–7697 (1999)). Subsequently, Hartwig and co-workers reported alkane and arene borylations utilizing much more active Rh precatalysts, such as Cp*Rh(η-C5Me5) (7)(Chen et al., Science 287: 1995–1997 (2000)).

A comparison of precatalysts 2 and 7 in borylations of various substituted arenes revealed that the Ir system was more selective towards arene C—H activation (Cho et al., J. Am. Chem. Soc. 122: 12868–12869 (2000)). Given the importance of selectivity in chemical synthesis, these findings spurred a detailed investigation of the original Ir system. Those results are described herein.

Compound 2 is stable in benzene solutions after prolonged thermolysis, which eliminates several mechanistic possibilities, including PMc dissociation to generate Cp*Ir (H)(BPin), an analog of proposed intermediates in the Rh system. However, added PMc strongly inhibits catalysis where HBPin is present. This indicated that small quantities of phosphine-Ir⁺ species could be active. Since Cp*IrH₄(BPin), species (x=1, 2) form in the thermolysis of Cp*IrH₄ and HBPin (Kawamura and Hartwig, J. Am. Chem. Soc. 123: 8422–8423 (2001), anisole borylations with identical loadings of Cp*IrH₄ and 2 were compared. From this experiment, Cp*IrH₄(BPin), intermediates can be eliminated because they are not kinetically competent for catalysis and the borylation regioselectivities for Cp*IrH₄ and 2 differ substantially. At 150° C., the following isomer ratios

\[
\begin{align*}
\text{Cp*IrH₄(BPin)} & \quad \text{Cp*IrH₄} \\
\text{Cp*IrH₄} & \quad \text{Cp*IrH₄(BPin)}
\end{align*}
\]
were obtained for anisole borylation with 20 mol % pre-catalyst loadings: Cp*IrH$_4$, o:m:p=3:49:48; 2 o:m:p=2:79:19.

Exclusion of a simple phosphine dissociative pathway narrows the plausible catalysts to two choices: (i) Ir phosphine species arising from Cp* loss or (ii) species where both Cp* and PMe$_3$ have been lost. The latter possibility was intriguing in light of Martin’s synthesis of (η$^5$-arene)Ir (BCat)$_2$ complexes (Cat=ortho-catecholate) from (Ind)Ir(COD) (3, Ind η$^5$-C$_5$H$_5$, COD=1,5-cyclooctadiene) and HBCat in arene solvents (10). Using an analogous route, we prepared (η$^5$-mesitylene)Ir(BPin)$_2$ (9 in which B(OH)$_2$ is BPin) in 19% yield from (Ind)Ir(COD) (8) and HBPin (Compound 9 has been prepared as an analytically pure white solid. Relevant spectroscopic data included $^3$H NMR (C$_6$D$_6$) δ 1.33 (s, 5H, BO$_2$C$_6$H$_3$), 2.23 (s, 9H, C$_5$H$_3$ - (CH$_3$)$_3$), 5.62 (s, 3H, C$_5$H$_3$(CH$_3$)$_3$). $^{11}$B NMR (C$_6$D$_6$) δ 32.5. $^{13}$C NMR (C$_6$D$_6$) δ 19.68, 25.73, 80.95, 96.97, 118.05). Compound 9 reacted with benzene at 150º C. to produce Ir metal and three equivalents of C$_6$H$_5$BPin, but did not catalyze C$_6$H$_5$BPin formation from benzene and HBPin. Thus, it appears that phosphines or related donor ligands are required for catalysis.

Utilizing the lability of the mesitylene ligand in 9, Ir phosphine species were generated in situ from 9 and appropriate phosphines and subsequently screened for activity. Borylation using 2 mol % 9 and 4 mol % PMe$_3$ was viable (Table 1, entry 1), and both catalytic activity and TONs for benzene borylation increased dramatically relative to precatalyst 2. Borylation rates were appreciable when P:Ir<3:1, but decreased dramatically when P:Ir ratio equals or exceeds 3:1.

**TABLE 1**

<table>
<thead>
<tr>
<th>Ent.</th>
<th>Sub.</th>
<th>Arene: HBPin</th>
<th>Prod.</th>
<th>Cat.</th>
<th>Ligand</th>
<th>Temp (ºC)</th>
<th>Time (h)</th>
<th>Yield (%)</th>
</tr>
</thead>
<tbody>
<tr>
<td>1</td>
<td>C$_6$H$_5$</td>
<td>16:1</td>
<td>PhBPin</td>
<td>(Me$_2$)Ir(BPin)$_3$ (9)</td>
<td>PMe$_3$</td>
<td>150</td>
<td>15</td>
<td>98b</td>
</tr>
<tr>
<td>2</td>
<td>C$_6$H$_5$</td>
<td>16:1</td>
<td>PhBPin</td>
<td>(Ind)Ir(COD) (8)</td>
<td>PMe$_3$</td>
<td>150</td>
<td>18</td>
<td>88b</td>
</tr>
<tr>
<td>3</td>
<td>C$_6$H$_5$</td>
<td>16:1</td>
<td>PhBPin</td>
<td>8</td>
<td>dppe</td>
<td>150</td>
<td>8</td>
<td>74b</td>
</tr>
<tr>
<td>4</td>
<td>C$_6$H$_5$</td>
<td>16:1</td>
<td>PhBPin</td>
<td>0.02 mol % 8</td>
<td>dppe</td>
<td>150</td>
<td>61</td>
<td>90b</td>
</tr>
<tr>
<td>5</td>
<td>C$_6$H$_5$</td>
<td>16:1</td>
<td>PhBPin</td>
<td>(IrCl(COD))$_2$</td>
<td>dppe</td>
<td>150</td>
<td>4</td>
<td>63</td>
</tr>
<tr>
<td>6</td>
<td>C$_6$H$_5$</td>
<td>16:1</td>
<td>PhBPin</td>
<td>8</td>
<td>dppe</td>
<td>150</td>
<td>3</td>
<td>81</td>
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<tr>
<td>10</td>
<td>Cl</td>
<td>1:1.5</td>
<td>PhBPin</td>
<td>8</td>
<td>dppe</td>
<td>150</td>
<td>14</td>
<td>89</td>
</tr>
<tr>
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<td>PhBPin</td>
<td>8</td>
<td>dppe</td>
<td>150</td>
<td>17</td>
<td>92</td>
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</table>
TABLE 1*-continued

<table>
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<tr>
<th>Ent.</th>
<th>Sub.</th>
<th>Arene: HBPin</th>
<th>Prod.</th>
<th>Cat.</th>
<th>Ligand</th>
<th>Temp (°C)</th>
<th>Time (h)</th>
<th>Yield (%)</th>
</tr>
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<tbody>
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<td>12°</td>
<td>Cl</td>
<td>N</td>
<td>Cl</td>
<td>8</td>
<td>dppe</td>
<td>100</td>
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<td></td>
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<td>BP in</td>
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<tr>
<td>13</td>
<td></td>
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<td>—</td>
<td>8</td>
<td>dppe</td>
<td>100</td>
<td>60</td>
<td>—</td>
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<td>9</td>
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<tr>
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<td>25</td>
<td>95</td>
</tr>
<tr>
<td></td>
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<td>MeO₂C</td>
<td>BP in</td>
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<td></td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>16</td>
<td></td>
<td>1:3</td>
<td>MeO</td>
<td>8</td>
<td>dmpe</td>
<td>150</td>
<td>95</td>
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<tr>
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<td></td>
<td>MeO</td>
<td>MeO</td>
<td></td>
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<td></td>
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</tbody>
</table>

*Reactions run in neat arene, Ir = 2 mol %, P: Ir = 2:1, and yields are reported for isolated materials unless otherwise noted. (COD = 1,5-cyclooctadiene, dmpe = Me₂PCH₂CH₂PMe₂, dppe = Ph₂PCH₂CH₂PPh₂).

GC yield based on HBPin.

Reactions run in cyclohexane.

The low isolated yields of 9 harpered screening efforts and precluded practical applications despite the dramatic improvement in catalytic activity. Hence, we sought alternative means for generating active catalysts. Since NMR spectra indicated quantitative generation of 9 from 8, in situ generation of active catalysts by phosphine addition to 8 was examined. Compound 8 was synthesized in 86% yield from indeny lithium and (Ir(COD))₂ (Merola and Kaemarik, Organometallics 8: 778–784 (1989)). This approach was successful and results for benzene borylations are shown in Table 1 (entries 2–5). Chelating phosphines substantially increased activity and TONs as highlighted for 1,2-bis(dimethylphosphino)ethane (dmpe) where the effective TON of 4500 (entry 5) represented an improvement of more than 1000-fold over precatalyst 2. In addition, active catalysts were generated from commercially available sources such as (Ir(COD))₂ (entry 6).

If the primary active species generated by PMMe₃ addition to 8 and 9 are identical to those generated from 2, borylation of substituted benzenes should exhibit similar regio- and chemoselectivities. Anisole is a useful substrate for probing regioselectivity and the meta:para ratios determined from borylations by active species generated by PMMe₃ addition to 8 and 9 are similar to those for 2 (For catalysts generated from 4 mol % PMMe₃ and 2 mol % 8 or 9, the following isomer ratios were obtained for anisole borylation at 150°C: 8, 8:9:9:74:17; 9, 8:9:9:8:75:17. For 8 and 9, ortho borylation increases slightly, which could signify a minor pathway that is not accessible from 2).

To assess chemoselectivities, the ratios of arene to benzylic activation in M-xylene were examined. The selectivities of catalysts generated from 8 (13:1) and 9 (12:1) were diminished relative to the selectivity of precatalyst 2 (35:1). Nevertheless, the Ir catalysts were more selective for arene activation than the Rh catalyst, 7, where the selectivity was 7:1 (Cho et al., J. Am. Chem. Soc. 122: 12868–12869 (2000)); a Rh catalyst that is highly selective for benzylic borylation has been recently reported (Shimada et al., Angew. Chem., Int. Ed. 40: 2168–2171 (2001)), and the addition of one equivalent of the chelating phosphine, 1,2-bis(diphenylphosphino)ethane (dppe) per equivalent of 8 or 9 generated catalysts where the arene to benzylic selectivities exceeded 142:1.

Dramatic differences in chemoselectivities between Ir and Rh catalysts were found for halogenated substrates, where the Ir catalysts preferentially activated C—H bonds. A representative procedure for borylation is given for entry 10 of Table 1. Briefly, in a glove box under N₂, compound 8 (57 mg, 0.14 mmol) and dppe (54 mg, 0.14 mmol) were dissolved in HBPin (1.30 g, 10.2 mmol). The solution was transferred to a thick-walled air-free flask containing 1,3-dichlorobenzene (1.00 g, 6.80 mmol). The clear yellow solution was heated at 100°C under N₂ and monitored by GC-FID. After 14 hours, the reaction mixture was pumped down to obtain a brown oil, which was vacuum distilled at 93–94°C (0.03 mmHg). The resulting oil was then dissolved in E₂O (10 mL) and washed with water (5×100 mL).
After drying over MgSO₄, ether was removed under high-vacuum to give 1.65 g (89% yield) of colorless 1,3,5-
C₃H₅Cl,BPin (mp 36-38°C). ¹H NMR (500 MHz, CDCl₃) δ 1.71 (t, J=2.0 Hz, 2H), 1.34 (d, J=2.0 Hz, 2H), 1.34 (s, J=2.0 Hz, 3H). ¹³C NMR (125 MHz, CDCl₃) δ 24.82, 84.0, 131.7, 133.4. ¹¹B NMR (CDCl₃) δ 30. Good yields of mono- or tri-borylated products of 1,3,5-triﬂuorobenzene were obtained by adjusting the acene:HPin ratio (Table 1, entries 7 and 8). In contrast, previous attempts to effect multiple borylations of 1,3,5-triﬂuorobenzene using the Rh catalyst 7 led to increased deﬂuorination (Cho et al., J. Am. Chem. Soc. 122: 12868-12869 (2000)). Borylations of aromatics with heavier halogen substituents provided an even starker contrast between Ir and Rh catalysts. For example, Ir catalyzed borylations of 1,3-dichlorobenzene and 1,3-difluorobenzene generate meta functionalized products in high yields (entries 10 and 11), while dehalogenation is the dominant pathway when the Rh catalyzed reaction (entry 12). Dehalogenation was observed during attempted silylations of 1,3-
dichlorobenzene using closely related Rh catalysts (Ezbiansky et al., Organometallics 17: 1455-1457 (1998)). The finding that aromatic C-halogen bonds survived in the Ir catalyzed reactions contrasted the Pd-catalyzed reactions of boranes and aryl bromides where the C—Br bonds were converted to C—B or C—H bonds (Murata, et al., J. Org. Chem. 65: 164-168 (2000)). Entry 12 illustrates an extension of meta selective borylation to a halogenated heterocycle.

Since ary1 iodides have the weakest carbon-halogen bonds, they are most susceptible towards reductive cleavage by transition metals. Hence, it is not surprising that the Ir catalysts generated from 8 were ineffective in aromatic borylation of iodobenzene (Table 1, entry 13). However, iodobenzene and HPin reacted smoothly to yield a mixture of C₆H₅HPin isomers when active catalysts were generated from the Irᴵᴵ source, 9, and dppe (entry 14). Thus, Ir catalysts are compatible with the etire range of aryl halides. Furthermore, functional group tolerance that was previously found in Rh catalyzed borylations extends to Ir catalyzed reactions (viz., ester compatibility in entry 15) and Ir selectively borylates symmetrical, 1,2-substituted arenes at the 4-position (entry 16).

The remarkable selectivity of Ir borylation catalysts for aromatic C—H bonds suggested that Ir byproducts might not interfere in subsequent reactions of the aryloborin products. Thus, we envisaged one-vesel elaborations of arene C—H bonds where catalytic borylations are followed by other metal-catalyzed events in a catalytic cascade. For a Zr/Pd catalyzed route to substituted biphenyls and terphenyls (see Frid et al., J. Am. Chem. Soc. 121: 9469-9470 (1999)), the union of catalytic borylations and Miyaura-Suzuki cross-couplings for one-
vesel bimetal synthesis from C—H and C—X precursors was attempted. As shown in FIG. 3A, the biaryl product can be prepared in good yield from the in situ Pd-catalyzed cross-coupling of 3-bromotoluene with 1,3,5-C₃H₅Cl,BPin, generated by Ir/Pd borylation of 1,3-dichlorobenzene with HPin. An interesting extension of Ir/Pd tandem catalysis highlighting Ir compatibility with halogenated aromatics is shown in FIG. 3B. The specific target was a hyperbranched polyphenylene that Kim and Webster prepared via Pd catalyzed coupling of the bromo/boronic acid monomer, 1,3,5-C₃H₅Br,B Bá(OH)₂ (Kim and Webster, Macromol. 25: 5561-5572 (1992)). Using Ir/Pd tandem catalysis, material with nearly identical NMR (¹H, ¹³C) and GPC data was obtained from HPin and 1,3-dibromobenzene in a one-vesel reaction. For the material in

![FIG. 3B](image-url)  
**M₃≈6374 and M₃≈3460 as compared to the previously reported values of M₃≈7570 and M₃≈3820 (Kim and Webster, Macromol. 25: 5561-5572 (1992)).**

From a mechanistic standpoint, catalytic cycles involving oxidative addition/reductive elimination from Irᴵᴵ and/or Irᴵᴵ intermediates are consistent with the results herein. Within this context, we considered Ir and Irᴵᴵ boryl intermediates to be the most likely C—H activating species in the Irᴵᴵ and Irᴵᴵ cycles, respectively. Hence, the Ir and Ir—I boryl complexes, Ir(BPin)(PMe₃), and fac-Ir(BPin)(PMe₃), were prepared in order to evaluate their stoichiometric reactions with arenes.

Compounds Ir(BPin)(PMe₃), and fac-Ir(BPin)(PMe₃) have been fully characterized as shown by the following spectral data: Ir(BPin)(PMe₃), ¹H NMR (CD₂Cl₂, 25°C) δ 1.24 (s, 12H, BO₂C₃H₅), 1.58 (t, 36H, PC(CHO)₂); ¹³B NMR (CD₂Cl₂) δ 38. ³¹P [¹H] NMR (CD₂Cl₂) δ 57.5; fac-Ir(BPin)(PMe₃), ¹H NMR (CD₂Cl₂) δ 1.34 (s, 36H, BO₂C₃H₅), 1.52 (m, 27H, PC(CHO)₂); ¹³B NMR (CD₂Cl₂) δ 36. ³¹P [¹H] NMR (CD₂Cl₂) δ -64. In reactions with arenes, compounds Ir(BPin)(PMe₃), and fac-Ir(BPin)(PMe₃) both reacted cleanly with benzene to produce PhBPin and the corresponding hydride complexes shown below, which was consistent with the idea that Ir or Ir—I species can effect arene borylation; however, the arene products from stoichiometric reactions of Ir(BPin)(PMe₃), and fac-Ir(BPin)(PMe₃) with iodobenzene differed substantially.

![Ir(BPin)(PMe₃)](image-url)  
**Ir(BPin)(PMe₃) + C₆H₆ → [Ir(BPin)(PMe₃)] + C₆H₅BPin**

![fac-Ir(BPin)(PMe₃)](image-url)  
**fac-Ir(BPin)(PMe₃) + C₆H₆ → fac-Ir(BPin)(PMe₃) + C₆H₅BPin**

Specifically, compound Ir(BPin)(PMe₃) reacted rapidly with iodobenzene at room temperature, but isomers of C₆H₅I(BPin) were not detected, even after prolonged thermolysis. Conversely, thermolysis of fac-Ir(BPin)(PMe₃) in iodobenzene produced m- and p-C₆H₅I(BPin) in 54% yield, based on fac-Ir(BPin)(PMe₃). In addition to a 45% yield of PhBPin.

Since conversion rates in catalytic reactions plummet when P:I ratios equal or exceed 3:1, the observation that Ir(BPin)(PMe₃) and fac-Ir(BPin)(PMe₃) were not kinetically competent for catalysis was expected. However, this does not exclude the possibility that identical intermediates are generated in the stoichiometric and catalytic reactions. Instead, generation of appropriate intermediates under catalytic conditions could simply be more efﬁcient. Nevertheless, the stoichiometric transformations lend credence to either IR or IR—I species mediating C—H activations under catalytic conditions. The reactions of Ir(BPin)(PMe₃) and fac-Ir(BPin)(PMe₃) with iodobenzene have greater mechanistic implications. For example, the absence of C₆H₅I(BPin) products in thermolysis of Ir(BPin)(PMe₃) mirrored the failed attempt to borylate iodobenzene using the Irᴵ precatalyst 8 (entry 13). The reactivity of fac-Ir(BPin)(PMe₃) suggests that an IR—I intermediate may activate C—H bonds in the presence of C—I bonds, but the chemistry of Ir(BPin)(PMe₃) is more important because it essentially excludes the participation of IR—I species in the successful borylation of iodobenzene using the Ir—I precatalyst 9 (entry 14).

Although catalytic processes involving IR—I intermediates have not been categorically excluded, we presently prefer the simplified mechanism involving IR—I and IR—I intermediates in FIG. 4 for the following reasons: (i) the correlations
between stoichiometric and catalytic borylations of iodo-
benzene by Ir and Ir[4] argue against an Ir[4] mechanism, (ii) the catalytic inhibition when P:Ir ratios equal or exceed 3:1 and the slow borylation rates for the 18-electron Ir[4] complex fac-Ir(IPh)2(PMe2Ph)3 are consistent with the generation of a reactive 16-electron bisphosphine Ir[6] intermediate from an 18-electron bisphosphine Ir[2] resting state, (iii) since chelating phosphines generally inhibit phosphine dissociative pathways, the catalytic activity with chelating phosphines supports the viability of bisphosphine intermediates, and (iv) the 18-electron bisphosphine compound, Ir(PMe2Ph)2H2, is an effective precatalyst for borylation. A more definite characterization of the catalytic manifold is underway.

In summary, an investigation of the original Ir catalytic system, whose promising selectivities could not be practi-
cally implemented due to extremely low TOFs, has produced a family of efficient borylation catalysts with remarkable regio- and chemoselectivities. In addition to providing a direct route to aryl and heteroaryl boron compounds from boranes and azirines, the viability of a tandem catalytic cascade where the first step is an Ir catalyzed aromatic borylation has been demonstrated. We are optimis-
tic that extensions of these findings will have significant synthetic applications.

EXAMPLE 2

This example shows a two-step, one-vessel process for synthesis of the biaryl 3,5-bis(trifluoromethyl)biphenyl from 1,3-bis(trifluoromethyl)benzene, a borane (HPBPin), and iodo benzene. An arylboronate ester was produced using an Ir catalyst and the arylboronate was subsequently coupled to the iodo benzene using a palladium catalyst.

In a glove box, HBPBPin (448 mg, 3.50 mmol) was added to a mixture of 1,3-bis(trifluoromethyl)benzene (500 mg, 1.34 mmol), (COD)Ir(Indenyl) (19.5 mg, 0.047 mmol), and dppe (18.6 mg, 0.047 mmol) in a small air-free flask equipped with a stir bar. The flask was then sealed and heated at 100°C for 16 hours.

Afterwards, the reaction solution was cooled to room temperature and Pd2(dba)3 (42.8 mg, 0.047 mmol), P(Ph3)(28.3 mg, 0.140 mmol), iodo benzene (476 mg, 2.24 mmol), K2PO4 (744 mg, 3.50 mmol), and DME (10 mL) were added. The mixture was stirred at 80°C for 3 hours. The 3,5-bis(trifluoromethyl)biphenyl was obtained (68.2% yield) as a colorless oil. The identity of the 3,5-bis(trifluoromethyl)biphenyl was confirmed by comparison to the GC retention time and 1H NMR data to an authentic sample prepared from 3,5-bis(trifluoromethyl)phenyl pina-
colborane and iodo benzene. 1H NMR (CDCl3, 300 MHz) δ 8.00 (s, 2H), 7.93 (s, 1H), 7.85–7.58 (m, 2H), 7.53–7.42 (m, 3H).

EXAMPLE 3

This example shows a closed system for a two-step, one-vessel process for synthesis of a biphenyl from benzene, a borane (HPBPin), and a halogenated phenyl (PhI). An arylboronate ester was produced using an Ir catalyst and the arylboronate was subsequently coupled to the halogenated phenyl using a palladium catalyst.

In a glove box, benzene (437 mg, 5.60 mmol), Ir(COD) (Indenyl) (3.0 mg, 0.0070 mmol), dppe (2.8 mg, 0.0070 mmol), dodecane (internal standard, 11.5 mg, 0.0675 mmol), and HBPBPin (45 mg, 0.352 mmol) were added to a J Young tube equipped with a stir bar. The tube was then sealed, removed from the box, and heated at 100°C for 18 hours. A GC trace of the reaction mixture revealed PhBPBPin in 85.7% yield.

Next, the reaction solution was allowed to cool to room temperature and Pd2(PPh3)4 (8.1 mg, 0.0077 mmol), K2PO4 (112 mg, 0.528 mmol), PhI (72.7 mg, 0.356 mmol) and DME (2 mL) were added. Three freeze-pump-thaw cycles were performed to remove residual oxygen and the reaction mixture was heated at 80°C for two days. GC analysis showed 81.8% GC-yield of the biphenyl, 79.7% GC-conversion of the iodo benzene, and 79.1% conversion of the PhBPBPin.

EXAMPLE 4

This example shows an open system for a two-step, one-vessel process for synthesis of a biphenyl from benzene, a borane (HPBPin), and a halogenated phenyl (PhI). An arylboronate ester was produced using an Ir catalyst and the arylboronate was subsequently coupled to the halogenated phenyl using a palladium catalyst.

In a glove box, benzene (1.0 mL, 11.2 mmol), Ir(COD) (Indenyl) (3.0 mg, 0.0070 mmol), dppe (2.8 mg, 0.0070 mmol), dodecane (internal standard, 12.0 mg, 0.07 mmol), and HBPBPin (45 mg, 0.352 mmol) were added to a Schlenk tube equipped with a stir bar. The tube was then sealed, removed from the box, and heated at 100°C for 18 hours with constant stirring.

Next, the reaction solution was cooled to room temperature and Pd2(PPh3)4 (8.5 mg, 0.0074 mmol), K2PO4 (112 mg, 0.528 mmol), PhI (74.9 mg, 0.367 mmol) and DME (2 mL) were added. A GC trace of the reaction mixture revealed PhBPBPin in 78.9% yield. The reaction mixture was then degassed by purging with nitrogen and stirred at 90–95°C for 16.5 hours. A GC analysis showed a 97.5% yield of the biphenyl, 84.1% GC-conversion of the iodo benzene, and 90.7% conversion of the PhBPBPin.

EXAMPLE 5

This example shows an open system for a two-step, one-vessel process for synthesis of a biaryl from 1,3,
dichlorobenzene, a borane (HPBPin), and a 3-bromotoluene.

In a glove box, 1,3-dichlorobenzene, Ir(COD)(Indenyl) (2 mol %), dppe (2 mol %), and HBPBPin were added to a Schlenk tube equipped with a stir bar. The tube was then sealed, removed from the box, and heated at 100°C for 16 hours with constant stirring.

Next, the reaction solution was cooled to room temperature and Pd2(PPh3)4 (2 mol %), K2PO4, 3-bromotoluene and DME (2 mL) were added. The reaction mixture was incubated at 80°C for 17 hours. A GC analysis following the reaction showed an 80% yield of the biaryl from the 1,3-dichlorobenzene.

EXAMPLE 6

This example shows an open system for a two-step, one-vessel process for synthesis of a hyperbranched polyphenylene from 1,3-dibromobenzene.

In a glove box, 1,3-dibromobenzene, Ir(COD)(Indenyl) (2 mol %), dppe (2 mol %), and HBPBPin were added to a Schlenk tube equipped with a stir bar. The tube was then sealed, removed from the box, and heated at 100°C for 16 hours with constant stirring.

Next, the reaction solution was cooled to room temperature and Pd2(dba)2 (0.5 mol %), K2PO4, 3-bromotoluene, and PhP were added. The reaction mixture was incubated for 15 hours with refluxing xylene. The
hyperbranched polyphenylene was similar to that obtained by the process of Kim and Webster (Macromol. 25: 5561-5572 (1992)).

While the present invention is described herein with reference to illustrated embodiments, it should be understood that the invention is not limited thereto. Those having ordinary skill in the art and access to the teachings herein will recognize additional modifications and embodiments within the scope thereof. Therefore, the present invention is limited only by the claims attached herein.

We claim:
1. A process for producing a substitute aromatic compound in one vessel, which comprises:
   (a) reacting in a reaction mixture first aromatic organic compound selected from the group consisting of an aryl, a six membered heterocaromatic compound, and a five membered heteroaromatic compound with a borane selected from the group consisting of a borane with a B—H, B—B, and B—Si bond in the presence of a catalytically effective amount of an ionic or rhodium complex with three or more substituents, and an organic ligand selected from the group consisting of phosphorus, carbon, nitrogen, oxygen, and sulfur organic ligands to form an aromatic boron compound in the reaction mixture; and
   (b) reacting the aromatic boron compound in the reaction mixture with a second aromatic organic compound selected from the group consisting of halide, triflate, and nonaflate in the presence of a catalytically effective, amount of a metal catalyst and in the presence of the iodide or rhodium, in the reaction mixture from step (a), wherein an aromatic group of the second aromatic organic compound is coupled to an aromatic group of the first aromatic organic boron compound to produce the substituted aromatic compound.

2. The process of claim 1, wherein the three or more substituents excludes hydrogen.

3. The process of claim 1 wherein the aromatic complex with the ligand is selected from the group consisting of (Cp*)Ir(H)(Me$_2$P), (Cp*)Ir(H)(BPin)(Me$_2$P), (Cp*)Ir(H)(C$_5$H$_7$)(Me$_2$P), (Ind)Ir(dppe), (Ind)Ir(dppe), (Mes)Ir(BPin)(B(OR)$_2$)$_2$, (R$_2$)P$_2$, (Ind)Ir(BPin)$_2$, (Ind)Ir(BPin)$_2$, (C$_5$H$_5$)Ir(OR)$_2$, ((R$_2$)P)$_2$, (C$_5$H$_5$)Ir((R$_2$)P)$_2$, ((R$_2$)P)$_2$, (C$_5$H$_5$)Ir(OR)$_2$, where x is 0-4, wherein Cp* is 1,2,3,4,5-methylcyclopentadienyl, BPin is pinacolborane, Me is methyl, H is hydrogen, P is phosphorus, Ind is indenyl, COD is 1,5-cyclooctadiene, MesH is mesitylene, and wherein R, R$_1$, and R$_2$ are hydrogen, linear or branched alkyl containing 1 to 8 carbons, aryl, or a carbon in a cyclic structure.

4. The process of claim 1 wherein the iridium complex is (Ind)Ir(COD) wherein Ind is indenyl and COD is 1,5-cyclooctadiene.

5. The process of claim 1 wherein the iridium complex with the ligand is (Mes)Ir(BPin)$_2$, wherein MesH is mesitylene and BPin is pinacolborane.

6. The process of claim 1 wherein the iridium complex is (Ind)Ir(COD)$_2$ wherein COD is 1,5-cyclooctadiene.

7. The process of claim 1 wherein the rhodium complex with the ligand is selected from the group consisting of (Cp*)Rh(H)(Me$_2$P), (Cp*)Rh(H)(BPin)(Me$_2$P), (Cp*)Rh(H)(C$_5$H$_7$)(Me$_2$P), and (Cp*)Rh(hexamethylbenzene), wherein Cp* is 1,2,3,4,5-methylcyclopentadienyl, BPin is pinacolborane, Me is methyl, H is hydrogen, and P is phosphorus.

8. The process of claim 1 or 2 wherein the phosphorus organic ligand is selected from the group consisting of trimethyl phosphine (PM$_3$), 1,2-bis(dimethylphosphino) ethane (dmpe), Ph$_2$P, Cy$_2$P, and 1,2-bis(diphenylphosphino) ethane (dppe).

9. The process of claim 1 wherein the borane is a borane ester.

10. The process of claim 1, wherein the borane is pinacolborane.

11. The process of claim 1 wherein the metal is palladium.

12. The process of claim 1 wherein the metal catalyst complex is selected from (Pd(PPh)$_3$)$_2$, Pd(dba)$_2$/P(Bu)$_3$, PdCl$_2$(dppe), and Pd(OAc)$_2$-Cy$_2$P wherein P is phosphorus and Ph is phenyl, dba is dibenzylideneacetone, ‘Bu is tert-butyl, dppe is diphenylphosphinoethane.

13. A process for producing an organic substituted aryl or heteroaryl compound, which comprises:
   (a) reacting in a reaction mixture in a reaction vessel a first aromatic organic compound selected from the group consisting of an aryl, a six membered heterocaromatic compound, and a five membered heteroaromatic compound with a borane selected from the group consisting of a borane with a B—H, B—B, and B—Si bond in the presence of a catalytically effective amount of an ionic or rhodium complex with three or more substituents, and an organic ligand selected from the group consisting of phosphorus, carbon, nitrogen, oxygen, and sulfur organic ligands to form an aromatic boron compound in the reaction mixture; and
   (b) reacting the aromatic boron compound in the reaction mixture with a second aromatic organic compound selected from the group consisting of halide, triflate, and nonaflate in the presence of a catalytically effective, amount of a metal catalyst and in the presence of the iodide or rhodium, in the reaction mixture from step (a), wherein an aromatic group of the second aromatic organic compound is coupled to an aromatic group of the first aromatic organic boron compound to produce the substituted aromatic compound.

14. The process of claim 13 wherein the three or more substituents excludes hydrogen.

15. The process of claim 13 wherein the iridium complex with the ligand is selected from the group consisting of (Cp*)Ir(H)(Me$_2$P), (Cp*)Ir(H)(BPin)(Me$_2$P), (Cp*)Ir(H)(C$_5$H$_7$)(Me$_2$P), (Ind)Ir(dppe), (Mes)Ir(BPin)(B(OR)$_2$)$_2$, (R$_2$)P$_2$, (Ind)Ir(BPin)$_2$, (Ind)Ir(BPin)$_2$, (C$_5$H$_5$)Ir(OR)$_2$, ((R$_2$)P)$_2$, (C$_5$H$_5$)Ir((R$_2$)P)$_2$, ((R$_2$)P)$_2$, (C$_5$H$_5$)Ir(OR)$_2$, where x is 0-4, wherein Cp* is 1,2,3,4,5-methylcyclopentadienyl, BPin is pinacolborane, Me is methyl, H is hydrogen, P is phosphorus, Ind is indenyl, COD is 1,5-cyclooctadiene, MesH is mesitylene, and wherein R, R$_1$, and R$_2$ are hydrogen, linear or branched alkyl containing 1 to 8 carbons, aryl, or a carbon in a cyclic structure.

16. The process of claim 13 wherein the iridium complex is (Ind)Ir(COD) wherein Ind is indenyl and COD is 1,5-cyclooctadiene.

17. The process of claim 13 wherein the iridium complex with the ligand is (Mes)Ir(BPin)$_2$, wherein MesH is mesitylene and BPin is pinacolborane.

18. The process of claim 13 wherein the iridium complex is (Ind)Ir(BPin)$_2$ wherein COD is 1,5-cyclooctadiene.

19. The process of claim 13 wherein the rhodium complex is selected from the group consisting of (Cp*)Rh(H)$_2$, (Me$_2$P), (Cp*)Rh(H)(BPin)(Me$_2$P), (Cp*)Rh(H)(C$_5$H$_7$)(Me$_2$P), and (Cp*)Rh(hexamethylbenzene), wherein Cp* is 1,2,3,4,5-methylcyclopentadienyl, BPin is pinacolborane, Me is methyl, H is hydrogen, and P is phosphorus.
20. The process of claim 13 wherein the phosphorus organic ligand is selected from the group consisting of trimethyl phosphine (PMe₃), 1,2-bis(dimethylphosphino)ethane (dmpe), Ph₂P, Cy₂P, and 1,2-bis(diphenylphosphino)ethane (dppe).

21. The process of claim 13 wherein the borane is a borane ester.

22. The process of claim 13 wherein the borane is pinacolborane.

23. The process of claim 13 wherein the metal is palladium.

24. The process of claim 13 wherein the metal catalyst complex is selected from Pd(PPh₃)₄, Pd₂(dba)₃/P(Bu)₃, PdCl₂(dppf), and Pd(OAc)₂/PCy₃ wherein P is phosphorus and Ph is phenyl, dba is dibenzylideneacetone, 'Bu is tert-butyl, dppf is diphenylphosphinoferrocene.

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